NAME: **HONORS CHEMISTRY**

SECTION: Boyle’s Law Problems

Standard Temperature and Pressure (STP): 0oC, 1 atm (or equivalent)

 760 mm Hg = 1 atm = 101.3 kPa = 760 torr

Remember to follow the general strategy:

Equation:

 List what you know

 Set up the problem

 Estimate and calculate

1. 100.0 mL of oxygen gas was collected at a pressure of 10.50 kPa. What is the gas volume if the pressure is changed to 9.91 kPa at constant temperature?

1. 50.0 cm3 of hydrogen was collected at a pressure of 97.3 kPa. If the pressure is changed to 101000 Pa, what is the new volume of the gas, assuming no temperature change?
2. What is the volume of sulfur dioxide at 101.3 kPa if 500.0 cm3 was collected at 95.6 kPa? Assume the temperature was constant.
3. A weather balloon containing 2.00 m3 of nitrogen was prepared at a pressure of 98.2 kPa. The balloon was released and came to a final volume of 2.50 m3. What was the final pressure of the nitrogen?
4. A flask containing 90.0 cm3 of hydrogen was collected under a pressure of 97.5 kPa. At what pressure would the volume be 70.0 cm3, assuming the temperature is kept constant?
5. A gas has a volume of 275 cm3 when measured at a pressure of 9.80 x 104 Pa. If the temperature is not changed, what would the gas volume be at standard pressure?
6. A gas has a volume of 50.0 m3 at standard pressure. Assuming no temperature change, what volume will the gas occupy
	1. if the pressure is doubled?
	2. If the pressure is tripled?
	3. If the original pressure is cut in half?
7. What is the volume occupied by 10.0 L of gas at standard pressure after it has been compressed at constant temperature to 500.0 kPa?
8. A gas is confined in a cylinder with a movable piston at one end. When the volume of the cylinder is 760.0 mL, the pressure of the gas is 125.0 kPa. When the cylinder volume is reduced to 450.0 mL, what is the pressure?

Answers

1. 106 cm3 O2
2. 48.2 cm3 H2
3. 472 cm3 SO2
4. 78.6 kPa
5. 125 kPa
6. 266 cm3
7. a) 25.0 m3

b) 16.7 m3

c)1.00 x 102 m3

8. 2.03 L

9. 211.1 kPa