NAME: **HONORS CHEMISTRY**

SECTION: Pairs/Check: Writing Names and Formulas

1. You will work with a partner for this activity. The younger partner will do the even problems, and the older partner will do the odd problems. After completing 2 problems each, trade papers and check your partner’s work. If a problem is correct, mark it with a check. Circle any incorrect answers. If you cannot agree on an answer, ask your teacher for assistance.
2. Complete the following table. Consider the elements that make up the compound. A few are shown as examples. Be sure to include Roman numerals for Type II cations and parentheses as needed. (Remember, the group 1 metals, group 2 metals, aluminum, gallium, zinc and silver form only one ion!) Also, remember that prefixes are ONLY used for molecular substances (Type III binary compounds) and that acids are named according to whether they are binary acids or oxyacids.

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| **Name** | **Ionic?** | **Molecular?** | **Acid?** | **Formula** |
| Sulfurous acid |  |  | x | H2SO3 |
| Tungsten(IV) hydroxide | x |  |  | W(OH)4 |
| Disulfur trioxide |  | x |  | S2O3 |
| 1. Hydroiodic acid |  |  |  |  |
|  |  |  |  | N2S3 |
|  |  |  |  | VCO3 |
| 1. Manganese(II) sulfide |  |  |  |  |
| 1. Sulfur pentabromide |  |  |  |  |
|  |  |  |  | HNO2(aq) |
| 1. Palladium(III) cyanide |  |  |  |  |
|  |  |  |  | PBr6 |
| 1. Magnesium hydrogen carbonate |  |  |  |  |
|  |  |  |  | Ag3P |
|  |  |  |  | Au2Cr2O7 |
| 1. Chromic acid |  |  |  |  |
| 1. Aluminum sulfite |  |  |  |  |
|  |  |  |  | HC2H3O2(aq) |
| 1. Zinc chromate |  |  |  |  |
|  |  |  |  | CF4 |
|  |  |  |  | H3PO4(aq) |
| 1. Niobium(IV) nitrate |  |  |  |  |
| 1. Carbon monoxide |  |  |  |  |
| 1. Dinitrogen tetrasulfide |  |  |  |  |
|  |  |  |  | Co(MnO4)3 |
| 1. Ammonium hydrogen phosphate |  |  |  |  |